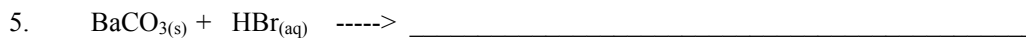
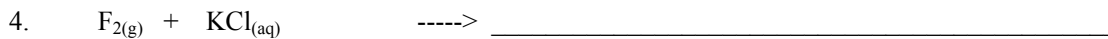
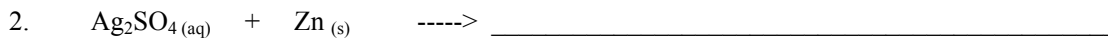
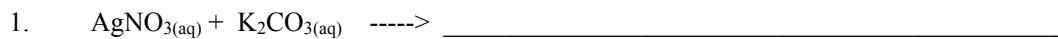
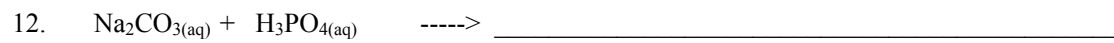
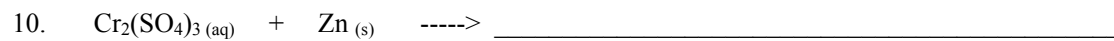
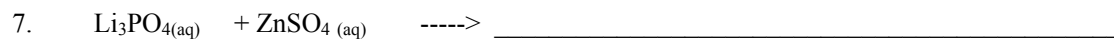


Net Ionic Equation Practice:

Predict products (including phase subscripts) and balance each reaction.

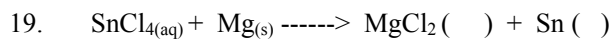
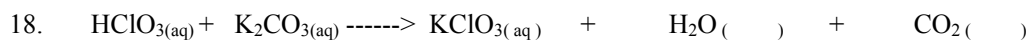
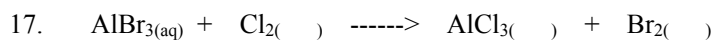
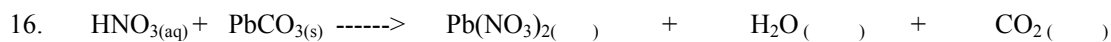
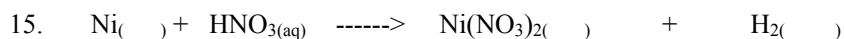
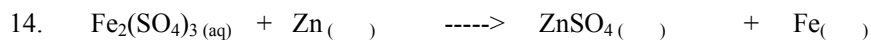
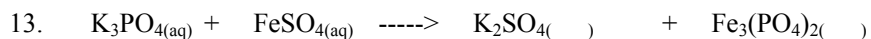
Then write a balanced net ionic equation, with phase subscripts, for each reaction.





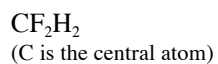
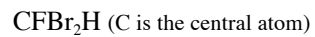
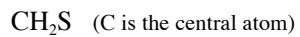
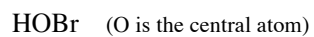
More Net Ionic Equation Practice:

For each equation: balance it, add missing subscripts, and write a balanced net ionic equation, with subscripts.



20. In reactions 13-19, state what (if anything) is oxidized, what is reduced, and the total number of electrons transferred in your balanced net ionic equation. Also identify any spectator ions.

21. Draw the Lewis Dot structure for each molecule or ion.



22. For each of the following compounds:

Draw the structure. Include charges on each ion.

Label at least one ionic bond and at least one covalent bond (if applicable) in the drawing.

If any of the ions include covalent bonds, draw the complete dot structure of the ion, including lone pairs, for at least one of the ions in the drawing.

