We will have a quiz on _____.
Topics: Lewis Dot Structures, Reactions and net ionic equations, and Oxidation and Reduction.
Study WS 8.8, 15.1, 16.0B, 16.1B, 16.2B, and the Qualitative Analysis Lab #2 (days 1 and 2)

Net Ionic Equation Practice:
Predict products (including phase subscripts) and balance each reaction.
Then write a balanced net ionic equation, with phase subscripts, for each reaction.

1. $AgNO_{3(aq)} + K_2CO_{3(aq)}$ -----> _____

2. $Ag_2SO_{4\,(aq)} + Zn_{\,(s)}$ -----> _____

3. $Ba_{(s)} + H_2O_{(l)}$ -----> _____

4. $F_{2(g)} + KCl_{(aq)}$ -----> _____

5. $BaCO_{3(s)} + HBr_{(aq)}$ -----> _____

6. $Cd_{(s)} + HNO_{3(aq)}$ -----> _____

7. $Li_3PO_{4(aq)} + ZnSO_{4\,(aq)}$ -----> _____

8. $Pb(OH)_{2(s)} + HNO_{3(aq)}$ -----> _____

9. $H_2O_{(l)}$ + $Na_{(s)}$ -----> _____

10. $Cr_2(SO_4)_{3\,(aq)}$ + $Zn_{\,(s)}$ -----> _____

11. $AgNO_{3\,(aq)}$ + $Al_{\,(s)}$ -----> _____

12. $Na_2CO_{3(aq)}$ + $H_3PO_{4(aq)}$ -----> _____

More Net Ionic Equation Practice:

For each equation: balance it, add missing subscripts, and write a balanced net ionic equation, with subscripts.

13. $K_3PO_{4(aq)}$ + $FeSO_{4(aq)}$ -----> $K_2SO_{4(\ \ )}$ + $Fe_3(PO_4)_{2(\ \ )}$

14. $Fe_2(SO_4)_{3\,(aq)}$ + $Zn_{(\ \ )}$ -----> $ZnSO_{4\,(\ \ )}$ + $Fe_{(\ \ )}$

15. $Ni_{(\ \ )}$ + $HNO_{3(aq)}$ ------> $Ni(NO_3)_{2(\ \ )}$ + $H_{2(\ \ )}$

16. $HNO_{3(aq)}$ + $PbCO_{3(s)}$ ------> $Pb(NO_3)_{2(\ \ )}$ + $H_2O_{(\ \ )}$ + $CO_{2\,(\ \ )}$

17. $AlBr_{3(aq)} + Cl_{2(\ \ )} \text{------>} AlCl_{3(\ \ )} + Br_{2(\ \ )}$

18. $HClO_{3(aq)} + K_2CO_{3(aq)} \text{------>} KClO_{3(\ aq\ )} + H_2O_{(\ \ )} + CO_{2\ (\ \ )}$

19. $SnCl_{4(aq)} + Mg_{(s)} \text{------>} MgCl_2\ (\ \ ) + Sn\ (\ \ )$

20. In reactions 13-19, state what (if anything) is oxidized, what is reduced, and the total number of electrons transferred in your balanced net ionic equation. Also identify any spectator ions.

21. Draw the Lewis Dot structure for each molecule or ion. (many of these are repeated from WS 16.1-16.3)

PCl_3 SO_3 HOBr (O is the central atom)

CH_2S (C is the central atom) $CFBr_2H$ (C is the central atom)

CO_3^{-2} IO_3^{-1} IO_2^{-1}

CF_2H_2
(C is the central atom) SO_2 NO_2^{+1}

OF_2 OH^{-1} NH_3

PO_4^{-3} NO_3^{-1} SO_3^{-2}

CSe_2 SSe_2 SI_2

CN^{-1} H_2 F_2 O_2 N_2